

A new series of ferromagnetic substances: the
ferrites of the rare earths

H. Forestier and G. Guiot-Guillain

Comptes Rendus de l'Academie des Sciences, Paris
230, 21 (22 May, 1950) 1844-1845

(from French)

Summary. Preparation of a new series of ferromagnetic compounds of the ferrites type, corresponding to the general formula $Fe_2O_3 \cdot M_2O_3$, in which M represents one of the elements of the rare earths; study of their stability and the variation of their magnetisation as a function of the temperature.

The general method of coprecipitation made it possible for one of us⁽¹⁾ to prepare in the pure state a series of compounds of the type $Fe_2O_3 \cdot MO$ (ferrites) and to demonstrate their magnetic properties. We adopted this method (frequently made use of since then) to attempt to obtain a new series of ferromagnetic substances of the type $Fe_2O_3 \cdot M_2O_3$, where M stands for one of the rare earths La Pr Nd Sm Er and Y.

The reaction $Fe_2O_3 + M_2O_3 = Fe_2O_3 \cdot M_2O_3$ is brought about by reheating of the coprecipitated mixture of the corresponding oxides. They are stable at high temperature ($1000^\circ C$). $Fe_2O_3 \cdot Nd_2O_3$, however is decomposed by heating for several hours at 1000° , into Fe_2O_3 and Nd_2O_3 (identifiable by examination in X-rays and the Curie point of Fe_2O_3); here an analogy may be seen with certain unstable ferrites which we had already investigated⁽²⁾.

The Curie Points occur for these ferrites of rare earths, at the following temperatures:

Fe_2O_3	{	La_2O_3	-	465°	Fe_2O_3	{	Sm_2O_3	-	300°
		Pr_2O_3	-	425			Er_2O_3	-	255
		Nd_2O_3	-	300			Y_2O_3	-	275

They are classified as follows, in descending order of magnetisation: Nd; Er; Y; Sm; Pr; La; magnetisation determined after heating above the Curie point followed by cooling in a field of 2000 gauss. For comparison we quote that nickel ferrites has, under the same experimental conditions, a magnetisation four times as great as $Fe_2O_3 \cdot Nd_2O_3$. We here reproduce the curves of thermomagnetic analysis, obtained by means of our recording apparatus.

On the other hand, ferrites of lanthanum and praseodimium show to a marked degree, the phenomenon of thermoremanent magnetism, discovered by one of us⁽³⁾ in sesquioxide of rhombohedral ferric oxide and the ferrites $Fe_2O_3 \cdot MO$. This magnetisation reaches a value 20 times greater than the initial magnetisation of Fe_2O_3 , La_2O_3 and 10 times greater than that of $Fe_2O_3 \cdot Pr_2O_3$.

Except in the case of neodymium ferrites, we observed that the proportion of ferromagnetic ferrites formed (in general very small at a temperature of 700° , even after prolonged heating) remained

/less

less than 50% after prolonged heating at 900°. In this way we obtained, for example, with $\text{Fe}_2\text{O}_3 \cdot \text{Fe}_2\text{O}_3$ after heating for 6 hours at 1000°, a product of which the magnetisation was three times as great as after heating for the same length of time at 920°. Such a high temperature was not necessary for the formation of the ferrites $\text{Fe}_2\text{O}_3 \cdot \text{MO}$ (completely formed after heating for four hours at 900°).

On the other hand, for $\text{Fe}_2\text{O}_3 \cdot \text{Nd}_2\text{O}_3$ (unstable ferrites), the maximum magnetisation (which can be attributed to about 80% of the product formed, according to the X-ray spectrograms) was obtained by heating for three hours at 775°.

All these results demonstrate the existence of a new series of ferromagnetic substances, of which the crystalline structures are being investigated at present.

BIBLIOGRAPHY

- (1) H. FORESTIER Thesis Paris 1908
- (2) H. FORESTIER and) C.R. de l'Acad.Sci. 193 (1931) 733.
M. GALAND)
H. FORESTIER and) C.R. de l'Acad.Sci. 199 (1934) 270.
G. GUIOT-GUILLAIN)
H. FORESTIER and) C.R. de l'Acad.Sci. 203 (1936) 1160
F. REDSLOB)
- (3) H. FORESTIER Ann. de Chim. 10th series 9 (1928) 389
H. FORESTIER and) C.R. de l'Acad. Sci. 183 (1926) 787
G. CHAUDRON)
H. FORESTIER C.R. de l'Acad. Sci. 201 (1935) 45
A. MICHEL and) C.R. de l'Acad. Sci. 200 (1935) 2171
G. CHAUDRON)

